Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Flow Diagram for carbonate chemistry

Air

Oceans

Sea floor

**Step 1 – pH of water**

Measure out about 100 mL of tap water into an Erlenmeyer flask. Add a few drops of pH indicator. What’s the pH? Is the solution acidic or basic?

**Step 2 – Effect of CO2**

Bubble some CO2 through the water in your flask. What’s the pH now? The reaction you just carried out can be expressed as

followed by

which ***releases*** , and therefore makes the solution **more acidic** (Did it?). Circle these reactions and label them as “Step 2” on your flow diagram.

**Step 3 – Effect of carbonate ion () on pH**

Add a little sodium carbonate to your flask What’s the pH now? The reaction you just carried out can be expressed as

which ***consumes*** , and therefore makes the solution less acidic (Did it?). Circle this reaction and label it as “Step 3” on your flow diagram.

**Step 4 – Properties of calcium carbonate ()**

Empty out your flask, and (just like Step 1) add water and a few drops of pH indicator. Now add a little calcium carbonate to your flask. What’s the pH now? How does it compare to the pH of ocean water (pH = 8.2, and falling)? Does it look like the carbonate is fully dissolving? The reaction you just carried out is

Circle this reaction and label it as “Step 4” on your flow diagram.

**Step 5 – A super-acidified ocean**

Add some concentrated hydrochloric acid to your flask. What’s the pH now? Does it look like the calcium carbonate is dissolving?